

WINE SHELF LIFE IMPROVEMENT

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- •What is Wine Shelf life?
- The Redox chemistry of the wine
 - Redox potential
 - Oxidation reactions
- Managing oxidation reactions during winemaking process
- Innovative tools to evaluate wine resistance to oxidation and manage oxygen
- •Q&A



THE RIGHT PRODUCT AT THE RIGHT TIME





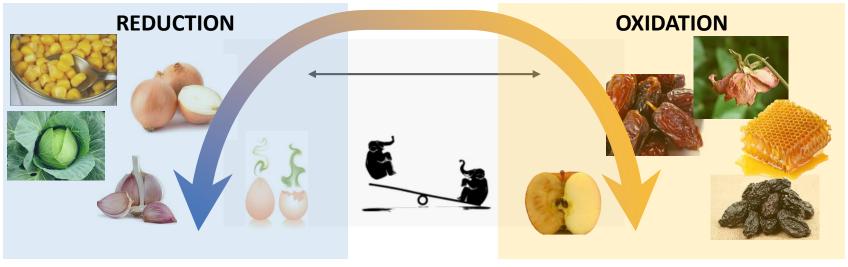






WINE EVOLUTION AND LONGEVITY



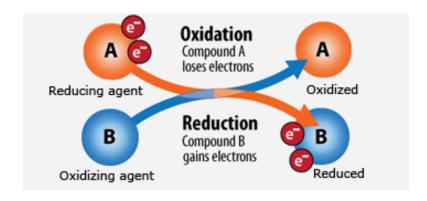




REDOX CHEMISTRY

Transfer of electrons

- Oxidation = loss of electron
- Reduction = gain of electron



Redox potential (mV) = tendency to gain or yield

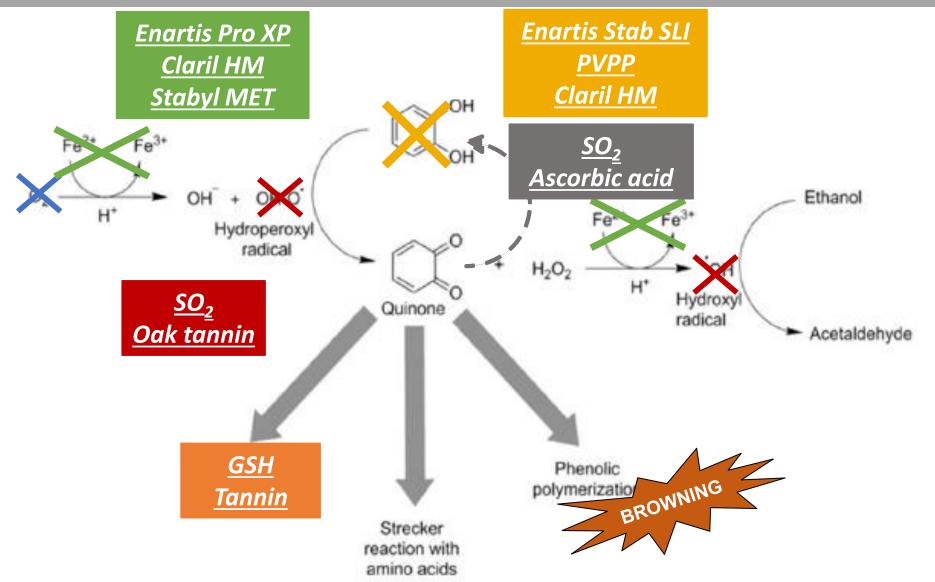
electrons

Reaction	Eh (mV)	Eh (mV)
	ph 3.5	ph 4
H2O2 + 2H + 2e = 2H2O	1570	1540
O2 + 4H + 4e = 2H2O	1020	990
Fe3+ + 1e = Fe2+	770	770
O2 + 2H + 2e = H2O2	490	460
Cu2+ + 1e = Cu+	160	160
S + 2H + 2e = H2S	-70	-100
Acetaldehyde + 2H + 2e = Ethanol	-410	-440

Metals
acids/pH
Phenolic compounds
Aldehydes
Ethanol
microorganisms
Aromatic compounds
Glutathion



OXIDO-REDUCTION REACTIONS





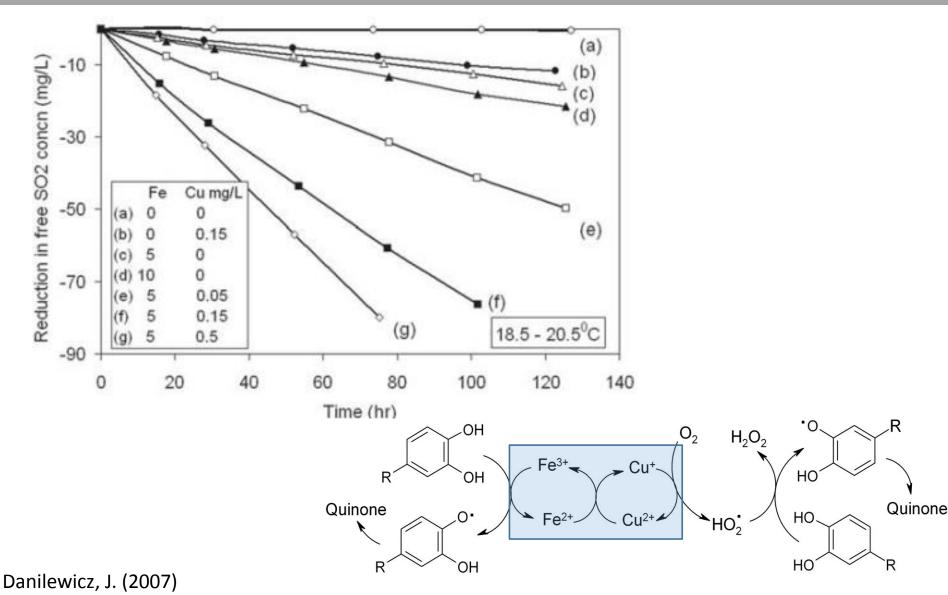


REMOVE CATALYZERS

Prevent oxidation
Improve wine redox stability
Any time of the wine life!

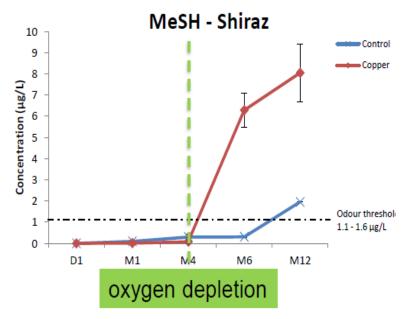


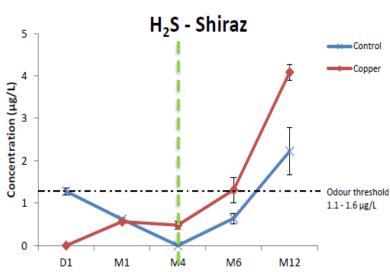
HOW IMPORTANT IS THE ROLE OF METALS IN WINE?

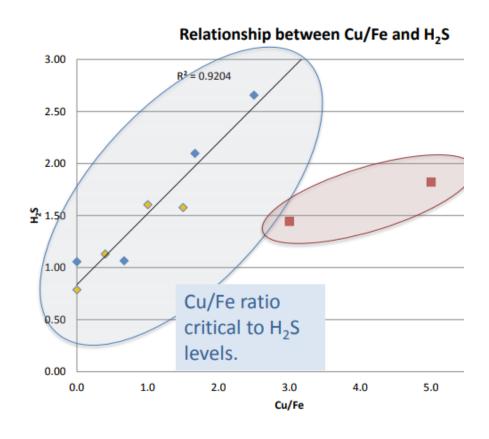




RESIDUAL COPPER AND VSC



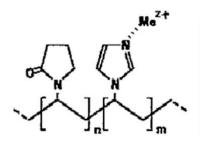




AWRI, Marlize Bekker, ASEV, 2017

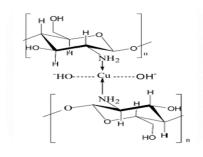


PVI/PVP AND CHITOSAN



PVI/PVP

$$Cu^{2+} > Au^{2+} = Ag^{2+} > Fe^{3+} > Mn^{2+} > Al^{3+}$$



Chitosan

$$Fe^{3+} > Cu^{2+} > Au^{2+} = Ag^{2+} > Mn^{2+} > Al^{3+}$$



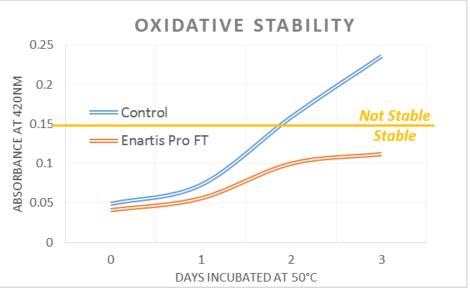




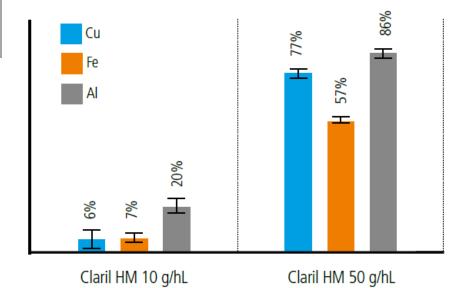


RESULTS...

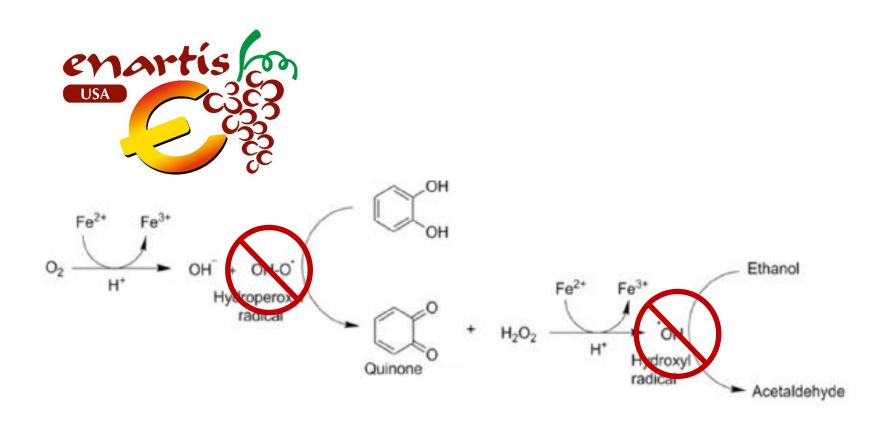




REDUCTION OF METALS IN WINE IN % COMPARING TO CONTROL







OXYGEN RADICAL SCAVENGING



TAN SLI: OXYGEN SCAVENGER



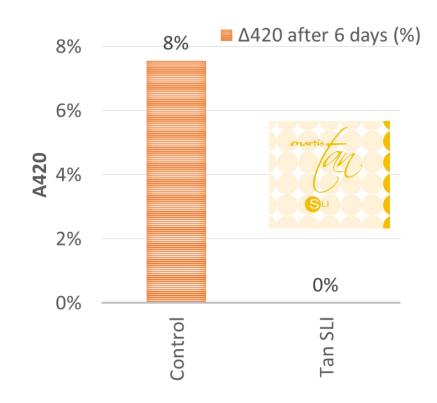
enartis —

- Scavenge radicals and limit oxidation
- Stabilize redox potential
- Binds with mercaptans to treat reduction

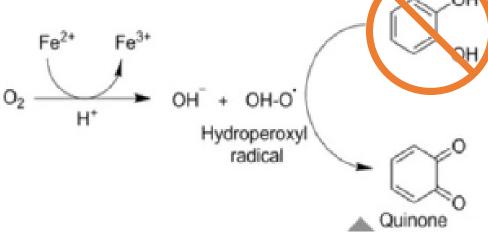
Applications:

- Transfer, racking
- Pre-bottling
- Treat reduction
- Extend wine shelf life

Δ420- 6 DAYS AFTER 2 RACK OFF







REDUCE SUBSTRATES/PRECURSORS

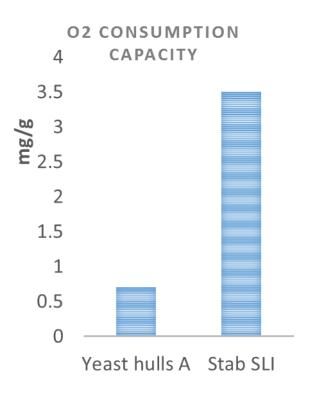
Phenolic compounds
Dissolved oxygen
Reduce VSC precursors

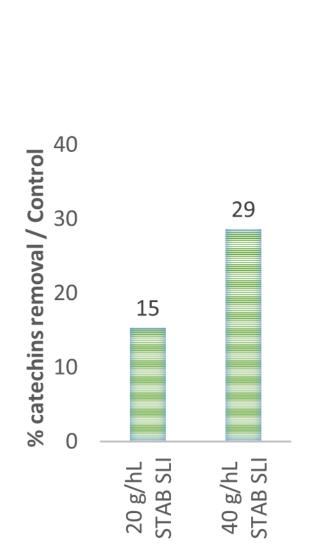


REDUCE OXIDATION SUBSTRATES

ENARTIS STAB SLI

- 'Active' lees
- PVPP
- Oak tannin







Wine after 6 months ageing on shelf. Control VS Stab SLI at 20 g/hL













Yeast metabolites

- Yeast nutrition
- SH amino acids

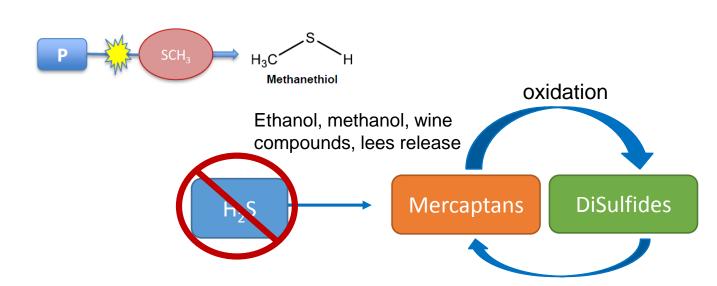
Precursors

- Elemental S
- SH amino acids
- Unknown



Oxygen management

- Yeast metabolism
- Quinone formation
- Oxidation of mercaptans to disulfides





 $O = O = \begin{bmatrix} H \\ S & O \end{bmatrix}$

SO₂ as a reductive agent : - 60 mV

Ascorbic acid as reductive agent: - 140 mV

REDUCING AGENT



PREPARATION FOR BOTTLING

Check your wine stability

- Browning test
- Oxidative stability
- Pinking test
- Antioxidant Capacity (RedOX/CaOX)



Oxygen management during bottling Choice of closure



TAKE HOME MESSAGES

Wine Shelf life: the right product at the right time

Redox potential is essential

Stabilize redox potential

- Eliminate metals
- Tan SLI

Limit precursors of oxidation and/or reduction

- Remove catechins with fining or Stab SLI
- Limit dissolved oxygen
- Yeast nutrition
- Limit reductive lees ageing

Check wine stability before bottling



INTERESTING LITERATURE

- Review of Oxidative Processes in Wine and Value of Reduction Potentials in Enology. Danilewicz, 2011
- •Interaction of Sulfur Dioxide, Polyphenols, and Oxygen in a Wine-Model System: Central Role of Iron and Copper. Danilewicz, 2007
- The Redox potential of Juice and wine. Boulton, 2017
- Controlling Redox Potential during fermentations. Boulton 2016
- •The effects of pH and copper on the formation of volatile sulfur compounds in Chardonnay and Shiraz wines post-bottling. Marlize Z. Bekker, 2016
- •Myths and facts regarding the role of precursors in the formation of 'reductive aroma' compounds in wines post-bottling. Marlize Z. Bekker, 2016
- The role of trace metals in wine 'reduction'. AWRI.





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THANK YOU FOR YOUR ATTENTION!